

The Classical Approximation

- *Idea:* From general properties we expect the interparticle potential $v(y)$ to have, and using a suitable approximation, we can draw some conclusions on the general behavior of $B(T)$ in some range of temperatures.
- *Setup:* For a classical gas,

$$\begin{aligned} Z_N &= \frac{1}{N! h^{3N}} \int d^3 p_1 \dots d^3 p_N \int d^3 r_1 \dots d^3 r_N e^{-\beta [\sum_i p_i^2 / 2m + V(\mathbf{r}_1, \dots, \mathbf{r}_N)]} \\ &= \frac{1}{\lambda^{3N} N!} \int d^3 r_1 \dots d^3 r_N e^{-\beta V(\mathbf{r}_1, \dots, \mathbf{r}_N)} = \frac{1}{\lambda^{3N} N!} \int d^3 r_1 \dots d^3 r_N e^{-\beta \sum_{i < j} v_{ij}} . \end{aligned}$$

- *Second virial coefficient:* From the general expression for B_2 , we find that classically (recall $Z_1 = V/\lambda^3$),

$$Z_2 = \frac{V}{2\lambda^6} \int d^3 y e^{-v_{12}(y)/k_B T} , \quad \text{and} \quad B_2 = -(Z_2 - \frac{1}{2} Z_1^2) \frac{V}{Z_1^2} = -\frac{1}{2} \int d^3 y f_{12}(y) ,$$

where $f_{12}(y) := e^{-v(y)/k_B T} - 1$ is a function that, contrary to $v(y)$, is well-behaved as $y \rightarrow 0$. Notice that this last expression for B_2 is the one that will be generalized in the cluster expansion to give all B_n .

- *Qualitative behavior of $B(T)$:* To proceed, we need to know something about $v(y)$. Let's assume that it has an infinite potential wall $v_{\text{HC}}(y)$ at $y = y_0$, and a weakly attractive form for $y > y_0$. Then (\star plot)

$$f(y) = \begin{cases} -1 & \text{for } |y| < y_0 \\ -v(y)/k_B T & \text{for } y > y_0 \end{cases} , \quad \text{and} \quad B(T) \approx -\frac{1}{2} \left[-\frac{4\pi}{3} y_0^3 + 4\pi \int_0^\infty dy y^2 \frac{v(y)}{k_B T} \right] = b - \frac{a}{k_B T} ,$$

where b is proportional to the molecular volume, and a arises from the attractive part of the potential.

Application to the Lennard-Jones Potential

- *Idea:* Use the phenomenological expression, that approximates well the measured potential for some gases,

$$v_{\text{LJ}}(r) = 4\epsilon \left[\left(\frac{\sigma}{r} \right)^{12} - \left(\frac{\sigma}{r} \right)^6 \right] .$$

- *Second virial coefficient:* Substituting into the general expression, and defining $r^* := r/\sigma$, $T^* := k_B T/\epsilon$,

$$\begin{aligned} B &= -\frac{1}{2} \int d^3 y (e^{-\beta v(y)} - 1) = \frac{2\pi}{3} \sigma^3 \frac{4}{T^*} \int dr^* r^{*2} \left(\frac{12}{r^{*12}} - \frac{6}{r^{*6}} \right) e^{-4(r^{*-12} - r^{*-6})/T^*} \\ &= \frac{2\pi}{3} \sigma^3 \left(\frac{1.73}{T^{*1/4}} - \frac{2.56}{T^{*3/4}} - \frac{0.87}{T^{*5/4}} - \dots \right) \end{aligned}$$

where we have integrated by parts in the second step.

Derivation of the van der Waals Equation of State

- *Idea:* We will apply a different kind of approximation, the mean-field approximation in which correlations between particles are ignored, and derive a common form for the equation of state for a dilute gas.
- *Setup:* Assume that the interparticle potential energy is of the form $v(y) = v_{\text{HC}}(y) + w(y)$, where the hard-core term v_{HC} is infinite for $y < y_0$, and w is regular. We want to get the equation of state from

$$Z = \frac{1}{\lambda^{3N} N!} \int d^3 x_1 \dots d^3 x_N e^{-\beta \sum_{i < j} v_{ij}} .$$

- *Mean-field approximation:* The mean field felt by the j -th particle due to the remaining $N - 1$ particles is

$$\sum_{i \neq j} w(y_{ij}) \approx (N - 1) \frac{1}{V} \int d^3 x w(x) = (N - 1) \bar{w} , \quad \text{so} \quad \sum_{j, i < j} w(y_{ij}) \approx \frac{1}{2} N^2 \bar{w} ,$$

and, if we define the parameter a by $\bar{w} =: 2a/V$,

$$Z \approx \frac{1}{\lambda^{3N} N!} \int_{y_{ij} > y_0} d^3 r_1 \dots d^3 r_N e^{-\beta N^2 \bar{w}/2} \approx \frac{(V - V_0)^N}{\lambda^{3N} N!} e^{N^2 a/V k_B T} ,$$

where $V_0 = (N - 1) b$, and $b := \frac{4}{3} \pi y_0^3$ is the effective volume of each molecule.

- *Equation of state:* From the free energy is $F = -k_B T \ln Z = -N k_B T \ln(V - V_0)/\lambda^3 + k_B T \ln N! - N^2 a/V$,

$$p = -\left. \frac{\partial F}{\partial V} \right|_{T, N} = \frac{N k_B T}{V - V_0} + \frac{N^2 a}{V^2} , \quad \text{or} \quad p = \frac{k_B T}{v - b} - \frac{a}{v^2} \quad (v := V/N) .$$

Relevant Sections: Phys 731; Chandler; Halley; Reif; Schwabl.